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# Acid Base

**chapter 15: acids and bases acids and bases** - either the brønsted-lowry acid, the brønsted-lowry base, the conjugate acid, or the conjugate base. identify the conjugate acid-base pairs in the reaction.  $\text{H}_2\text{SO}_4(\text{aq}) + \text{HPO}_4^{2-}(\text{aq}) \rightleftharpoons \text{HSO}_4^-(\text{aq}) + \text{H}_2\text{PO}_4^-(\text{aq})$  strong vs. weak acids and bases strong acids: strong electrolytes - completely ionized in solution there are 6 strong ... **acid-base chemistry - myresourceoerix** - a stronger acid has a weaker conjugate base, with water having a higher affinity for the hydrogen ion than the weaker conjugate base. on the other hand, a weaker acid has a stronger conjugate base, with the stronger conjugate base having a higher affinity for the hydrogen ion than water. **experiment 3: acid/base extraction and separation of ...** - benzoic acid use of two different bases with two different strengths allows for selective reaction of the stronger acid versus the weaker acid. the weaker base, sodium bicarbonate, is strong enough to react with the stronger acid, benzoic acid, but not strong enough to react with the weaker acid, 2-naphthol. **chapter 12 { acid-base chemistry - webassign** - a lewis acid is a substance that has an empty orbital that it can use to share a lone pair to form a bond. a lewis base is a substance that has a lone pair that it can share in a covalent bond. a lewis acid-base reaction is the conversion of the lone pair on the base and the empty orbital of the acid into a covalent bond between the acid and ... **test 2 ch17a acid-base practice problems** - b. the weaker its conjugate base. e. the more concentrated the conjugate base. c. the more concentrated the acid. 11. ammonia ( $\text{NH}_3$ ) acts as a weak base in aqueous solution. what is the acid that reacts with this base when ammonia is dissolved in water? a. none, there are no acids in pure water b.  $\text{H}_2\text{O}$  c.  $\text{NH}_4^+$  d. **intro to acids & bases worksheet** - base is a substance that when added to water increases the concentration of  $\text{OH}^-$  ions present.  $\text{HCl}$  is an example of an arrhenius acid and  $\text{NaOH}$  is an example of an arrhenius base.  $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$   $\text{NaOH} \rightarrow \text{Na}^+ + \text{OH}^-$  classify the following as an acid or base using the arrhenius definition of an acid or base. 7. **hno 3 fluid, electrolytes, acid-base and shock** - 4. discuss common causes of fluid, electrolyte and acid-base imbalances 5. discuss nursing responsibilities for assessment, identification, treatment and follow-up of fluid, electrolyte and acid-base imbalances 6. discuss nursing responsibilities for hypovolemic, cardiogenic, neurogenic, anaphylactic and septic shock . readings: 1. **experiment 7 - acid-base titrations** - an acid/base neutralization reaction will yield salt and water. in an acid-base titration, the neutralization reaction between the acid and base can be measured with either a color indicator or a ph meter. acid + base salt + water

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